

Review Sheet Chapter 18 Reaction Rates/Equilibrium Mrs Sanford Name:

Rate-

In Chemistry rate is expressed as-

Collision theory: In order for a successful reaction to take place, atoms have to collide with enough_____.

Activation energy-

Activated complex-

Another name for the activated complex is-

The factors that affect reaction rate are: (write in table below)

Factor	How does it affect rate?

Inhibitor-

Sketch a labelled activation energy diagram for an EXOthermic reaction

Sketch a labeled activation energy diagram for anENDOthermic reaction

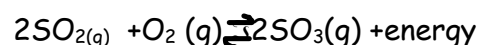
Give an ex. Of a reversible reaction with the proper arrows-

Chemical Equilibrium-

Dynamic equilibrium means-

Le Chatelier's principle:

How do the following affect each equilibrium?



Concentration		
Temperature		
Pressure		
Catalyst		

Keq=

If $K_{eq} > 1$ then _____	If $K_{eq} < 1$ then _____	
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K _{sp} -
The smaller the K _{sp} value the lower the _____.
Common ion-
Common ion effect-
If $aA + bB \rightarrow cC + dD$ the rate expression is-
How do you decide order of a reactant?
How do you determine overall order?
Reaction mechanism-
Intermediate-
Reaction steps are also known as _____ reactions.
Do p 555 #6 p 557 #7,8 p 558 #9, p 562 17, 18 p 564 #20 p 581 51a, 52b,c, 53b, 67, p 582 75, 76,78,79,81, 82, p 583 83